**Study Guide: Stoichiometry**

Vocabulary you should know:

Actual Yield

Balanced Chemical Equation

Composition Stoichiometry

Excess Reactant

Ideal Stoichiometric Calculation

Limiting Reactant

Molar Mass

Mole

Mole Ratio

Percent Yield

Products

Reactants

Reaction Stoichiometry

Stoichiometry

Theoretical Yield

* Distinguish between ideal and real stoichiometric calculations.
* Distinguish between limiting reactants and excess reactants in a chemical reaction
* Be able to determine amount of excess reactant used and remaining
* Know how the value of the theoretical yield generally compares with the value of the actual yield
* Be able to explain why actual yields are usually less than calculated theoretical yields
* Be able to explain the concept of mole ratio as used in reaction stoichiometry problems
* Know what the source of the mole ratio
* Be able to write and balance equations and know why a balanced equation is necessary to solve a mass-mass stoichiometry problem
* Be able to calculate percentage yield and what data are necessary to calculate the percent yield of a reaction

Review textbook pages in Chapter 9 on pp. 289- 308. Look over Chapter summary on page 310. Review modeling and practice problems, lab, notes, review, sample problems in textbook.

Look at my Resources page on my website for more sites that you can practice and check your work.

Remember this is just a guide! We have practiced, read, discussed, and covered all concepts on test in a variety of ways in and out of class.